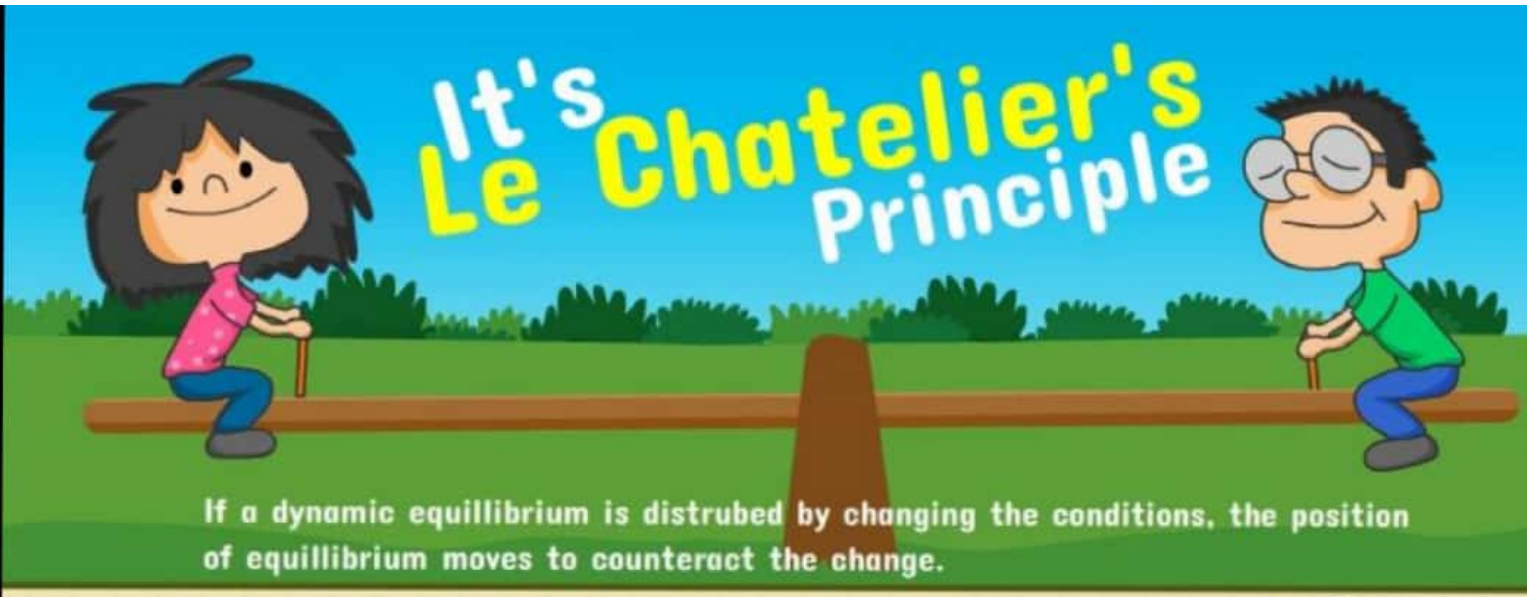
It's Le Chatelier's Principle



If a dynamic equillibrium is distrubed by changing the conditions, the position of equillibrium moves to counteract the change.

i will destroy your equilibrium

i will re-establish it

Change	Reactants			⇌	Products			Counteract
	a A	+	b B		c C	+	d D	
I will increase reactant concentration				⟷				Then I will shift the reaction forward
I will steal products				⟷				I will shift the reaction forward
I will increase pressure				⟷				I will reduce number of moles
I will decrease pressure				⟷				Then I'll increase number of moles
I will heat up your exothermic reaction				⟷				I'll shift the reaction backward
I will put your endothermic reaction in ice				⟷				I'll warm it up by forward reaction
I will catalyze your reaction				⟷				Hahaha.... It won't disturb my equilibrium
I will add noble gases to your reaction				⟷				Hahaha.... It won't disturb my equilibrium

How did you bypass my tricks ? ⟷ It's Le Chatelier's principle dear !